

There are two ways of *citing references in text*: (1) by author name, (2) by number in order of appearance. Below are a couple examples. The colored text in the examples is only because these are screenshots of hyperlinks in a web page. Do not add color in your text. When you are writing for a journal, check the specific requirements of that journal's publisher. Also see these web pages on technical writing <<http://owl.english.purdue.edu/owl/section/4/16/>> and formatting <<http://owl.english.purdue.edu/owl/resource/747/1/>>.

## BY AUTHOR

from: O. Ashrafi I, N. Mostoufi, R. Sotudeh-Gharebagh , "Two phase steady-state particle size distribution in a gas-phase fluidized bed ethylene polymerization reactor," *Chemical Engineering Science*, v. 73, pp. 1–7 (2012).

Despite inherent importance of PSD, limited number of papers has been published on the modeling of the particle size distribution in gas-phase catalytic olefin polymerization processes. Zacca et al. (1996) developed the population balance model (PBM) using catalyst residence time to model particle size developments in multistage olefin polymerization reactors. Choi et al. (1994) followed the population balance approach of Kunii and Levenspiel (1991) to investigate the effect of feed catalyst size distribution on the PSD of the polyolefin produced in fluidized bed reactor for both non-deactivated and deactivated catalysts. Soares and Hamielec (1995) developed a mathematical model to analyze the effects of residence time in ideal and non-ideal reactors on the polymer particles PSD which was produced by Ziegler–Natta catalyst. They showed that the residence time has unique effect on PSD. Khang and Lee (1997) applied a population balance approach to investigate the effect of non-ideal mixing behavior of solid particles on PSD in the fluidized bed olefin polymerization reactor. Arastoopour et al. (1988) experimentally studied the agglomeration behavior of inactive polyethylene particles. Their findings indicated that the rate of formation of large agglomerates was exponentially dependent on the bed temperature. Hatzantonis et al. (1998) formulated and

## References

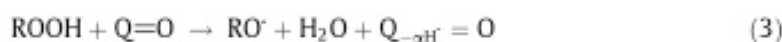
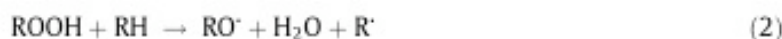
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## BY NUMBER

This involves a little more work because you need to list and number the references in the reference list at the end of the paper in the order that they are first cited in the text. This example puts the reference numbers in square brackets; another style is to put the number as a superscript.

from: U. Neuenschwander, I. Hermans, "Thermal and catalytic formation of radicals during autoxidation." *J. Catalysis*, v. 287, pp. 1-4 (2012)

C-centered radical, eventually stabilized by delocalization [5]. Such a substrate or product-induced initiation not only features a lower activation barrier, it is also proposed to be more efficient in generating radicals than the unimolecular O—O bond cleavage, because the nascent RO<sup>•</sup>-radical is effectively shielded by the initially hydrogen-bonded water molecule against recombination with the R<sup>•</sup>-radical<sup>1</sup>.



The RO<sup>•</sup> radicals are rapidly converted to R<sup>•</sup> radicals (reaction (4)) which are themselves trapped by O<sub>2</sub>, leading to ROO<sup>•</sup> radicals (reaction (5)). Peroxyl radicals are able to abstract H-atoms from the substrate and form ROOH (reaction (6)).



In addition, ROO<sup>•</sup> radicals can abstract the weakly bonded α H-atom of the ROOH primary product, leading to the formation of alcohol and ketone in an activated solvent-cage reaction (7); the stoichiometric coefficient x in reaction (7) depends on the substrate [6–9].

<sup>1</sup> H<sub>2</sub>O acts as "insulator" between the two radicals and is hydrogen-bonded to the alkoxy radical with approximately 2.6 kcal mol<sup>-1</sup>, at the ZPE-corrected UB3LYP/6-311++G(df,pd) level of theory.

## References

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