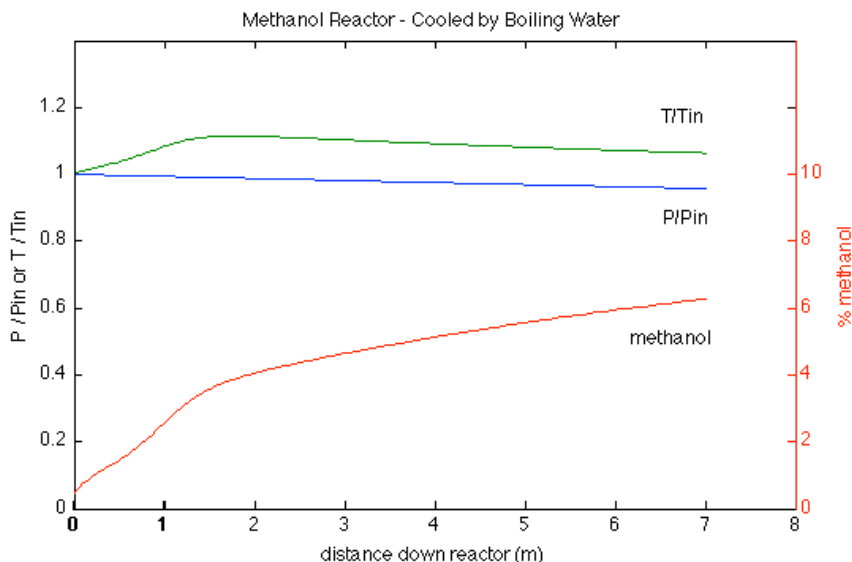


## Chemical Reaction Engineering - Part 13-A - Methanol Synthesis Reactor Model

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We wrote a Matlab program to simulate a methanol synthesis reactor. This is a 1-dimensional, plug flow model that assumes there are no radial gradients in the catalyst bed. The effectiveness factors of the catalyst are assumed constant. The reaction rates are from Vanden Bussche and Froment (1996).

The first figure shows some of the results of the model. Input values are from Chen et al. (2011) for a shell-and-tube reactor cooled by pressurized boiling water. The catalyst pellets are inside the tubes.



The methanol mole fraction at the reactor outlet is small such that reactants are separated from the methanol product and recycled to the reactor in a methanol plant. The methanol mole fraction is limited to low values by thermodynamic equilibrium constraints.

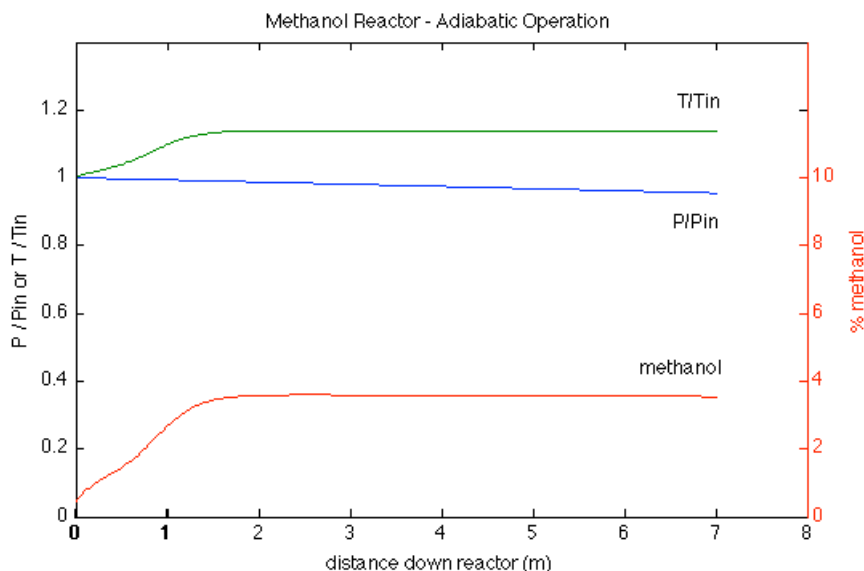
Chen et al. used the kinetics of Vanden Bussche and Froment in an Aspen Plus reactor model. They also had access to plant data for a large methanol reactor with 1620 each, 7-meter-long tubes. Below is a table comparing data from the plant, the results of their reactor model, and the results of this model.

	methanol mole fraction	T out (°C)	Pressure out (bar)
Chen et al. plant data	0.0630	255	66.7
Chen et al. model	0.0629	256	66.7
This model, $T_j = 231.2$ °C	0.0627	257	66.7
This model, $T_j = 230.4$ °C	0.0630	256	66.7

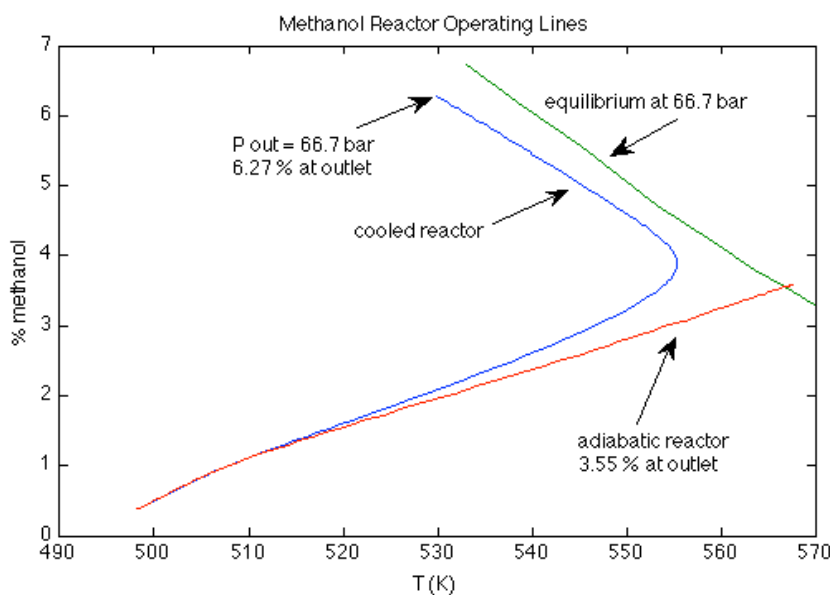
Although we used the same reaction kinetics as Chen et al., there are differences between the two models that would produce the somewhat different results. For example, they used a model of a counter-current heat exchange reactor in Aspen Plus for the boiling water (steam raising) reactor in the plant, whereas we used a single heat transfer jacket temperature.

In the first entry for this model in the table, the jacket temperature was set to 231.2 °C, which was the outlet steam temperature of Chen et al.'s reactor. The cooling water entered their reactor at 220 °C such that part of their tubes were exposed to a somewhat cooler temperature than the steam temperature. In the last entry for this model in the table, the constant jacket temperature was lowered by 0.8°C and closer agreement to the plant data was obtained.

The second figure shows the same reactor model except that it is run with adiabatic operation.



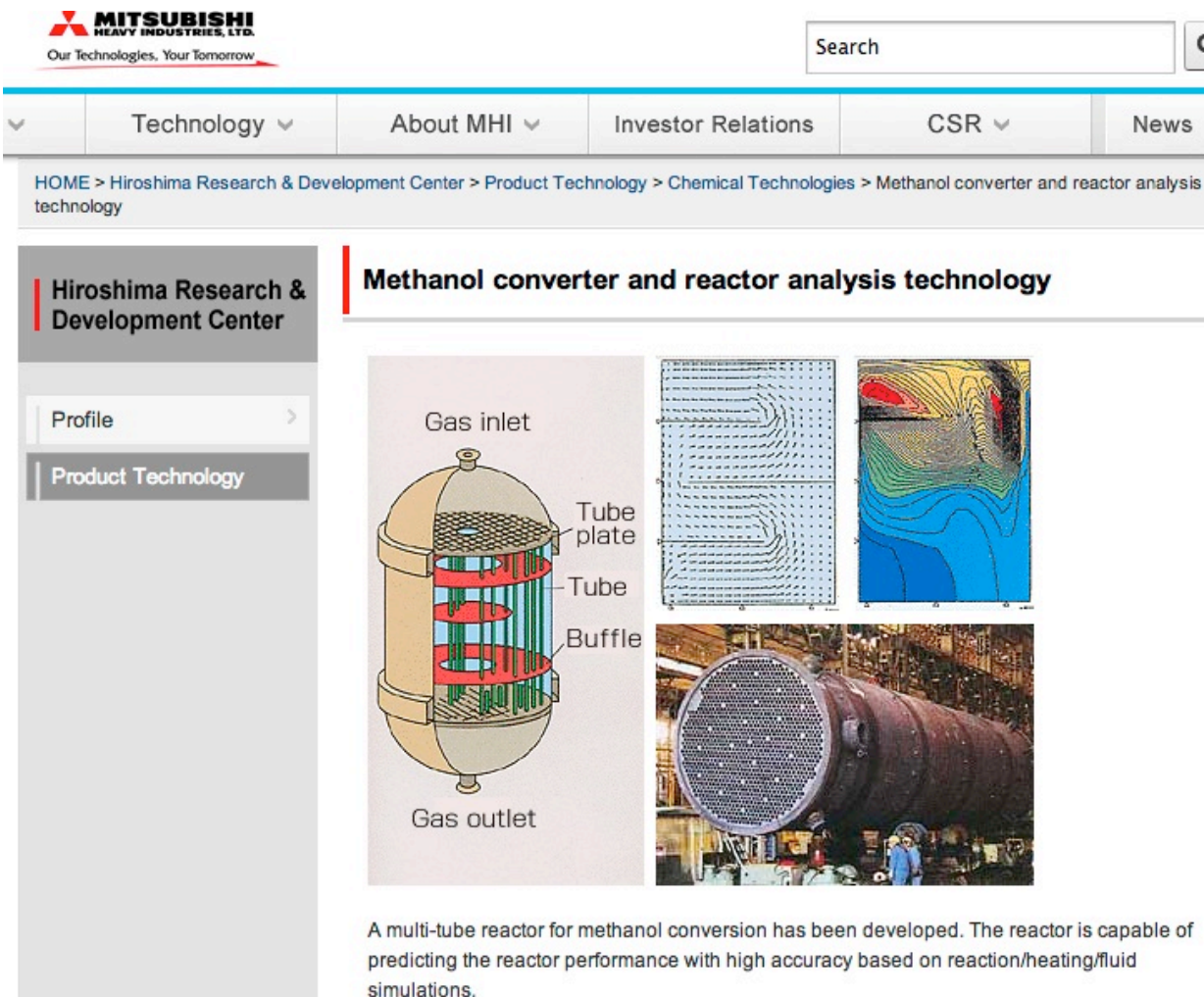
The third figure plots the operating lines of the cooled and adiabatic cases on a percent methanol vs. temperature plot. The equilibrium line at the outlet pressure of the cooled reactor case is also plotted. The adiabatic reactor operating line goes past the 66.7 bar equilibrium line because equilibrium is established in the middle of the adiabatic reactor where the pressure and temperature are higher than they are at the outlet, where equilibrium is also obtained (66.7 bar, 567 K, 3.55 % at outlet). The temperature falls near the end of the reactor as the pressure continues to fall because equilibrium is shifting to lower conversion to methanol and the reverse methanol synthesis reaction is endothermic.



## References

- K. M. Vanden Bussche and G. F. Froment, "A Steady-State Kinetic Model for Methanol Synthesis and the Water Gas Shift Reaction on a Commercial Cu/ZnO/Al<sub>2</sub>O<sub>3</sub> Catalyst," *J. Catalysis*, vol. 161, pp. 1-10, 1996.
- L. Chen, Q. Jiang, Z. Song, and D. Posarac, "Optimization of Methanol Yield from a Lurgi Reactor," *Chemical Engineering & Technology*, vol. 34, pp. 1-7, 2011.

The screenshot below shows a similar reactor. Note the people at the bottom of the photograph of the disassembled reactor. The catalyst is inside the tubes. The top right panels show that this company models the coolant flow profiles and temperature variations and in the shell. Our model assume a constant coolant temperature.



The screenshot displays the Mitsubishi Heavy Industries website. At the top left is the Mitsubishi Heavy Industries logo with the tagline "Our Technologies, Your Tomorrow". A search bar is located at the top right. Below the logo is a navigation menu with options: Technology, About MHI, Investor Relations, CSR, and News. The breadcrumb trail reads: HOME > Hiroshima Research & Development Center > Product Technology > Chemical Technologies > Methanol converter and reactor analysis technology.

The main content area is titled "Methanol converter and reactor analysis technology". On the left is a sidebar for the "Hiroshima Research & Development Center" with a "Profile" dropdown and "Product Technology" selected.

The central content features a 3D cutaway diagram of a multi-tube reactor. Labels include "Gas inlet" at the top, "Gas outlet" at the bottom, "Tube plate" at the top of the tube bundle, "Tube" for the individual tubes, and "Baffle" for the horizontal support structures. To the right of the diagram are two simulation plots: a vector field showing flow patterns and a contour plot showing temperature variations. Below these is a photograph of the physical reactor, a large horizontal cylindrical vessel, with several people standing at its base for scale.

A caption below the images reads: "A multi-tube reactor for methanol conversion has been developed. The reactor is capable of predicting the reactor performance with high accuracy based on reaction/heating/fluid simulations."

## Listing of main program MeOH.m - Listings of function files follow this

Read the first set of comments and note the structure of the main file and the user-written function files. The problem has been separated into modules to make it easier to write, debug, read and understand.

```
% Methanol synthesis reactor simulation
% by Richard K. Herz <rherz@ucsd.edu> 2011
% main MATLAB file:
%   MeOH.m
% other files required by the main file:
%   MeOHodes.m - ode's to integrate for T, P, stoichiometric extents
%   MeOHmolarflow.m - molar flow rates down bed
%   MeOHheatCp.m - gas mixture heat capacity
%   MeOHheatrxn.m - heats of reaction
%   MeOHviscosity.m - gas mixture viscosity
%   MeOHrates.m - reaction rates, currently from Vanden Bussche and Froment, 1996
%   MeOHenergy.m - dT/dW, temperature down bed
%   MeOHpress.m - dP/dW, pressure drop down bed

clear all
close all
fprintf('-----run separator-----\n')

% ----- SPECIFY INPUTS -----

% INPUT VALUES HERE ARE FROM Chen et al., 2011:
%   L. Chen, Q. Jiang, Z. Song, and D. Posarac,
%   "Optimization of Methanol Yield from a Lurgi Reactor,"
%   Chemical Engineering & Technology, vol. 34, pp. 1-7, 2011.

% INPUT INLET TEMPERATURE AND PRESSURE OF REACTANT GAS

Tin = 225 + 273.15; % K
Pin = 69.7; % bar

% INPUT REACTOR CATALYST BED PROPERTIES

Bdiam = 0.04; % m, ID of one catalyst tube
Blen = 7; % m, length of tubes
Btubes = 1620; % number of tubes

% NOTE: Chen et al., 2011 had cylindrical pellets 5.4 mm in diam x 5.2 mm long
Bdp = 0.0054; % m, diam of catalyst pellets
Bvoid = 0.285; % bed void fraction, voids between pellets
Bcatdens = 1190; % kg/m3, catalyst pellet density, where bed dens = cat dens*(1-Bvoid)

% INPUT HEAT TRANSFER DATA
%   for an adiabatic bed, set U = 0
%   to approach a true isothermal bed, set Tj = Tin and U = large value, e.g., 1e4 or 1e5
%   for a boiling water reactor (steam raising), enter appropriate Tj and U

Tj = 231.2 + 273.15; % K, heat transfer jacket temperature, Chen et al. steam temp 231.2 C
U = 118.44; % J/s/m2/K, heat transfer coefficient

% INDEPENDENT STOICHIOMETRIC EQUATIONS (ISE) DEFINING EXTENTS e()
%   ISE 1 for extent 1: CO2 + 3H2 = CH3OH + H2O
%   ISE 2 for extent 2: CO2 + H2 = CO + H2O

% initial values of extents
ein = [0, 0]; % mol/s

% INPUT INITIAL VALUES OF COMPONENT MOLAR FLOW RATES
%   order of components in vectors: H2, CO, CO2, MeOH, water, N2, CH4

% ONE WAY is to specify total molar flow rate and components as relative molar flows
%   and then compute vector of inlet molar flow rates, Fin

FtotIn = 1740.2; % mol/s, 1740.2 = 6264.8 kmol/h, 28.44 mol/s = 225.52 lbmol/h
```

```

% first list component flows as relative molar flow rates
% must have H2 and CO2 non-zero or will get div by zero in MeOHrates
% H2, CO, CO2, MeOH, water, N2 (inert), CH4 (inert)
xin = [9586.5/2 10727.9/28 23684.2/44 756.7/32 108.8/18 8072/28 4333.1/16];
% then compute mole fractions
xin = xin/sum(xin);
% then compute inlet molar flow rates
Fin = xin*FtotIn; % mol/s

% ALTERNATIVELY you can just specify vector of inlet molar flow rates, Fin
% must have H2 and CO2 non-zero or will get div by zero in MeOHrates

% ----- INTEGRATE DOWN REACTOR -----

% calculate total mass of catalyst
Wcat = Btubes*Bcatdens*(1-Bvoid)*pi*Blen*(Bdiam^2)/4 % kg, total mass catalyst

% Wspan = final catalyst weight (kg) for integration
Wspan = [0, Wcat];

% initial conditions for integration
yin = [Tin, Pin, ein];

% package bed and heat transfer conditions in an array
% to pass via ode45 to MeOHodes function
Bprop = [Bdiam,Btubes,Blen,Bdp,Bvoid,Bcatdens,Wcat,U,Tj];

% integrate using standard Matlab function ode45
[W, ya] = ode45('MeOHodes',Wspan,yin,[],Bprop,Fin);

% ----- POST PROCESSING -----

T = ya(:,1);
P = ya(:,2);
e1 = ya(:,3); % extent 1 for: CO2 + 3H2 = CH3OH + H2O
e2 = ya(:,4); % extent 2 for: CO2 + H2 = CO + H2O

% IDEAL GAS CONSTANT VALUES IF NEEDED FOR POST PROCESSING
R = 8.314472; % J/mol/K, for Arrhenius rate equations, ideal gas constant
Rg = 8.314472e-05; % bar-m3/mol/K, for ideal gas law, ideal gas constant

% can't use MeOHmolarflow here since e1 & e2 are column vectors

% molar flow rates in order: H2, CO, CO2, methanol, water, N2, CH4
% extent 1 for ISE 1: CO2 + 3H2 = CH3OH + H2O
% extent 2 for ISE 2: CO2 + H2 = CO + H2O
F(:,1) = Fin(1) - 3*e1 - e2; % H2
F(:,2) = Fin(2) + e2; % CO
F(:,3) = Fin(3) - e1 - e2; % CO2
F(:,4) = Fin(4) + e1; % methanol
F(:,5) = Fin(5) + e1 + e2; % water
F(:,6) = Fin(6); % N2 (inert)
F(:,7) = Fin(7); % CH4 (inert)

% get mole fractions
Ft = sum(F,2); % total
% F is a 2D array so can't divide by 1D Ft array
% need to construct a 2D array Ftot to get 2D array of mole fractions
Ftot = [Ft Ft Ft Ft Ft Ft Ft];
% mole fraction
x = F./Ftot;

ph2 = P.*x(:,1); % bar
pco = P.*x(:,2);
pco2 = P.*x(:,3);
pm = P.*x(:,4); % m = methanol
pw = P.*x(:,5); % w = H2O

xm = x(:,4); % mole fraction methanol

[r c] = size(P);

```

```

Pout = P(r)
if Pout < 1
    fprintf('WARNING: Pout < 1 bar. \n')
    fprintf('Increase number of tubes (area for flow), or \n')
    fprintf('decrease inlet flow rate. \n\n')
end
ToutK = T(r)
ToutC = T(r)-273.15
xMeOHout = x(r,4)

% mole fractions in order: H2, CO, CO2, methanol, water, N2, CH4
z = Blen*W/Wcat;

[AX,H1,H2] = plotyy([z z],[P/Pin T/Tin],z,100*xm); % , 'r--',z,T/Tin,'b--')
title('Methanol Reactor, left axis blue = P/Pin, green = T/Tin, right axis red = % methanol')
xlabel('distance down reactor (m)')
axis(AX(1),[0 8 0 1.4])
axis(AX(2),[0 8 0 14])
set(AX(1),'ytick',[0; 0.2; 0.4; 0.6; 0.8; 1.0; 1.2])
set(AX(2),'ytick',[0; 2; 4; 6; 8; 10])
set(get(AX(1),'ylabel'),'string','P / Pin or T / Tin')
set(get(AX(2),'ylabel'),'string','% methanol')

% from separate experiments to get equilibrium T & xMeOH at Pout = 66.7 bar
Teq = [569.9; 566.68; 563.63; 559.17; 556.2; 553.3; 545.21; 538.96; 533.11];
xmeq = [0.0330; 0.0355; 0.0380; 0.0419; 0.0445; 0.0472; 0.0558; 0.0613; 0.0672];

figure
plot(T,100*xm,'b',Teq,100*xmeq,'g')
title('Methanol Reactor, blue = with cooling, green = equilibrium at 66.7 bar')
ylabel('% methanol')
xlabel('T (K)')

```

## Listing of function file MeOHodes.m

```

function dydW = MeOHodes(W,y,flags,Bprop,Fin)
% by Richard K. Herz <rherz@ucsd.edu> 2011
% function MeOHodes is called by program MeOH
% uses function MeOHmolarflow
% uses function MeOHrates
% uses function MeOHenergy
% uses function MeOHpress
% inputs: y(1) = T, y(2) = P, y(3) = e(1), y(4) = e(2)
% outputs: dydW(1) = dT/dW, dydW(2) = dP/dW, dydW(3) = dFmeoh/dW, dydW(4) = dFco/dW
% T in K, P in bar
% r(1,1) = rate of methanol formation (mol/kg/s)
% r(2,1) = rate of CO formation (mol/kg/s)

T = y(1);
P = y(2);
e(1) = y(3);
e(2) = y(4);

% call 1st so can pass to other functions
F = MeOHmolarflow(e,Fin);

% call 1st so can pass rates to MeOHenergy function
r = MeOHrates(T,P,F); % mole/kg/s
r1 = r(1,1); % de(1)/dW = dFmeoh/dW
r2 = r(2,1); % de(2)/dW = dFco/dW

dydW(1,1) = MeOHenergy(T,P,F,Bprop,r1,r2); % dT/dW
dydW(2,1) = MeOHpress(T,P,F,Bprop,Fin); % dP/dW
dydW(3,1) = r1; % de(1)/dW = dFmeoh/dW
dydW(4,1) = r2; % de(2)/dW = dFco/dW

```

## Listing of function file MeOHheatCp.m

```
function Cp = MeOHheatCp(T,P,F)
% by Richard K. Herz <rherz@ucsd.edu> 2011
% is called by function MeOHenergy
% returns Cp of mixture (J/K/mol)

% need mole fractions x below
Ftot = sum(F);
x = F/Ftot;

% coefficients for fit of Cp to 3rd order polynomial in T from
% Handbook of Chemistry and Physics at 1.0133 bar
% rows in order: H2, CO, CO2, MeOH, Water, N2, CH4

Cpcoef(1,1:4) = [-1.0653e-09 5.3794e-06 -0.004198 30.1063];
Cpcoef(2,1:4) = [-2.8314e-09 6.0321e-06 0.002423 27.5047];
Cpcoef(3,1:4) = [8.1713e-09 -3.5920e-05 0.059157 22.9484];
Cpcoef(4,1:4) = [9.8419e-09 -5.7293e-05 0.12811 8.9427];
Cpcoef(5,1:4) = [-3.6093e-09 1.0553e-05 0.002698 31.7360];
Cpcoef(6,1:4) = [-3.6077e-09 9.0963e-06 -0.001325 28.4826];
Cpcoef(7,1:4) = [-2.3256e-09 -1.2497e-05 .076466 12.0092];

% adjust each component Cp for pressure
% from Cp's from Aspen Plus, 500-501 K & values at
% 1.0133 and 70 bar except as noted
Pfac = [5.2762e-5 % H2
        5.3402e-4 % CO
        1.6103e-3 % CO2
        6.9236e-4 % MeOH, 1.0133 & 10 bar
        1.3890e-3 % H2O, 1.0133 & 20 bar
        4.8512e-4 % N2
        6.1509e-4]; % CH4

% adjust to get agreement with Cp in Aspen at 500 K, 70 bar
% of Chen et al. (2011) reactor outlet mixture
Pfac = Pfac * 0.9523;

Pcorrec = 1 + (P - 1.0133)*Pfac;

Cp = 0;
for i = 1:7
    Cp = Cp + x(i) * Pcorrec(i) * polyval(Cpcoef(i,:),T);
end
```

## Listing of function file MeOHheatrxn.m

```
function Hrxn = MeOHheatrxn(T,P)
% by Richard K. Herz <rherz@ucsd.edu> 2011
% is called by function MeOHenergy
% returns delta-H reaction in (J/mol) for
% two Independent Stoichiometric Equations (ISE):
% ISE 1: CO2 + 3H2 = CH3OH + H2O
% ISE 2: CO2 + H2 = CO + H2O
% is called by function MeOHenergy
% get good agreement with Aspen Plus at all these conditions:
% 298 K and 1 bar
% 500 K and 1 bar
% 298 K and 60 bar
% 500 K and 70 bar

% need to correct del-Hf for both T and P

Hrxn298 = [-49.3160 ; 41.1540]; % kJ/mol, at 298.15 K, 1 atm

% correct Hrxn at 298.15 K and 1 atm for P
% from fit to Aspen results
% these pressure factors are for mixture of components
% specific to each stoich equation
Hrxn298(1) = Hrxn298(1) - 0.0686*(P - 1.0133); % P in bar
Hrxn298(2) = Hrxn298(2) - 0.0365*(P - 1.0133); % P in bar
```



```

% NOW NEED TO CORRECT FOR T NOT 298.15 K

% del-Hf(T) = del-Hf(Tr) + INTEG from Tr to T (del-Cp)

% delCpCoef = polynomial fit vs. T of
% Sum(stoich-coef-i * Cp-i) for each stoich equation
% for 3rd order fit, for Cp in J/mol/K
delCpCoef = [1.2571e-09  -2.6958e-05  0.084247  -72.5886
             -1.3547e-08  4.7125e-05  -0.049837  6.1860];
% now get integral of delCp over Tr = 298.15 to T
Tr = 298.15;
Tfac = [(T^4-Tr^4)/4 ; (T^3-Tr^3)/3 ; (T^2-Tr^2)/2 ; (T-Tr)];

% THIS Tfac CORRECTS HRXN AT 1 ATM FOR T OK
% but Cp's increase with P
% here adjust Tfac for each reaction separately for agreement
% with Aspen heat of reaction at 500 K & 70 bar,
% where these pressure factors are for specific mixture of
% components in each stoich equation
Tfac1 = Tfac * (1 - 0.00466*(P-1.0133)); % P in bar
Tfac2 = Tfac * (1 - 0.01550*(P-1.0133)); % P in bar

% now compute full T correction adjusted for P
delCpInt(1,:) = delCpCoef(1,:) * Tfac1; % J/mol
delCpInt(2,:) = delCpCoef(2,:) * Tfac2; % J/mol

% finally compute Hrxn corrected for T and P
Hrxn = Hrxn298 * 1e3 + delCpInt; % J/mol at T

```

## Listing of function file MeOHviscosity.m

```

function visc = MeOHviscosity(T,P,F)
% by Richard K. Herz <rherz@ucsd.edu> 2011
% is called by function MeOHpress
% returns gas dynamic viscosity in units of (Pa s)
% viscosity is corrected for T but not for pressure

% NOTE: this doesn't change much with composition so if need to
% speed up the program you can set to a constant average value
% for example, using viscosity for air gave results not too different
% from mixture calculations below
% Air viscosity from Physical Chemistry by P.W. Atkins (1978), p. 810
% mu = 1.82e-5; % kg/m/s, for air at 293 K
% mu = mu*(T/293).^0.5; % correct for T

% From Properties of Liquids and Gases, 5th ed,
% by Poling, Prausnitz & O'Connell
% at low P, dynamic viscosity is independent of P
% and proportional to T^(1/2)
% see Wilke method on p. 9.21 for gas mixtures

% From data over 300-600 K in CRC Handbook of Chem and Phys
% coefficients for gas dynamic viscosity for 1st order polynomial in T^0.5
% Viscosity is given in units of (micro-Pa s)
% rows: % H2, CO, CO2, methanol, water, N2, CH4
Viscoef(1,:) = [0.7526 -4.0884];
Viscoef(2,:) = [1.5750 -9.4437];
Viscoef(3,:) = [1.8116 -16.4474];
Viscoef(4,:) = [1.4234 -15.2870];
Viscoef(5,:) = [1.5913 -17.9877];
Viscoef(6,:) = [1.6324 -10.4033];
Viscoef(7,:) = [1.1434 -8.5867];

Thalf = sqrt(T);
Vis(1) = polyval(Viscoef(1,:),Thalf);
Vis(2) = polyval(Viscoef(2,:),Thalf);
Vis(3) = polyval(Viscoef(3,:),Thalf);
Vis(4) = polyval(Viscoef(4,:),Thalf);
Vis(5) = polyval(Viscoef(5,:),Thalf);
Vis(6) = polyval(Viscoef(6,:),Thalf);
Vis(7) = polyval(Viscoef(7,:),Thalf); % (micro-Pa s)

```



```

Vis = Vis * 1e-6; % (Pa s) = (micro-Pa s) * 1 Pa / 1e6 micro-Pa

% now need to get mixture viscosity for this composition

% From Properties of Liquids and Gases, 5th ed,
% by Poling, Prausnitz & O'Connell
% see Wilke method on p. 9.21 for gas mixtures

% mole fractions in order: H2, CO, CO2, methanol, water, N2, CH4
Ftot = sum(F);
x = F/Ftot;

MW = [2 28 44 32 18 28 16]; % g/mol, only be used in ratio so grams OK

for i = 1:7
    for j = 1:7
        tNum = ( 1 + (Vis(i)/Vis(j))^0.5 * (MW(i)/MW(j))^0.25 )^2;
        tDenom = ( 8*(1+MW(i)/MW(j)) )^0.5;
        phi(i,j) = tNum/tDenom;
    end
end

visc = 0;
for i = 1:7
    tDenom(i) = 0;
    for j = 1:7
        tDenom(i) = tDenom(i) + x(j)*phi(i,j);
    end
    visc = visc + x(i) .* Vis(i) ./ tDenom(i);
end

% HOWEVER, gas viscosity increases with P at high P
% so later should include this
% for now just multiply by a factor to get agreement with
% Chen et al. (2011) plant data
% visc = visc*(1 + 0.01146*(P - 1.0133));
visc = visc*(1 + 0.013*(P - 1.0133));

```

## Listing of function file MeOhrates.m

```

function r = MeOhrates(T,P,F)
% by Richard K. Herz <rherz@ucsd.edu> 2011
% is called by function MeOHodes
% returns rates of the two reactions in mol/s/kg
% inputs: y(1) = T, y(2) = P, y(3) = e(1), y(4) = e(2)
% T in K, P in bar
% e = stoichiometric extents
% r(1,1) = de(1)/dW = rate of methanol formation (mol/kg/s)
% r(2,1) = de(2)/dW = rate of CO formation (mol/kg/s)
%
% SEE COMMENTS BELOW TO ADD OR DELETE REACTION PATHWAYS

R = 8.314472; % J/mol/K, for use in Arrhenius rate expressions below, ideal gas constant

% need mole fractions x below
Ftot = sum(F);
x = F/Ftot;

% partial pressures
ph2 = P*x(1); % bar
pco = P*x(2);
pco2 = P*x(3);
pm = P*x(4); % m = methanol
pw = P*x(5); % w = H2O

% equilibrium constants for the two
% Independent Stoichiometric Equations (ISE)

% K1 for ISE 1: CO2 + 3H2 = CH3OH + H2O
log10K1 = 3066/T - 10.592;
K1 = 10^log10K1;

```

```

% K2 for ISE 2: CO2 + H2 = CO + H2O
log10K2 = -2073/T + 2.029;
K2 = 10^log10K2;

% rates of reaction pathways from Vanden Bussche and Froment, 1996
% K. M. Vanden Bussche and G. F. Froment,
% "A Steady-State Kinetic Model for Methanol Synthesis
% and the Water Gas Shift Reaction on a Commercial Cu/ZnO/Al2O3 Catalyst,"
% J. Catalysis, vol. 161, pp. 1-10, 1996.

% rate coefficient parameters from Table 2
A(1) = 0.499; % pre-exp for sqrt(Kh2)
B(1) = 17197; % energy for sqrt(Kh2)
A(2) = 6.62e-11; % pre-exp for Kw
B(2) = 124119; % energy for Kw
A(3) = 3453.38; % pre-exp for Kw/K8/K9/Kh2
B(3) = 0; % energy for Kw/K8/K9/Kh2
A(4) = 1.07; % pre-exp for k'5aK'2K3K4Kh2
B(4) = 36696; % energy for k'5aK'2K3K4Kh2
A(5) = 1.22e10; % pre-exp for k'1
B(5) = -94765; % energy for k'1

% rate coefficients
k = A .* exp(B/R/T);

denom = (1+k(3))*pw/ph2 + k(1)*sqrt(ph2) + k(2)*pw;

% R1 is the rate of methanol formation only in the
% reaction pathway CO2 + 3H2 = CH3OH + H2O
R1 = k(4)*pco2*ph2*(1 - pw*pm/ph2^3/pco2/K1)/denom^3; % mol/s/kg

% R2 is the rate of CO formation only in the
% reaction pathway CO2 + H2 = CO + H2O
R2 = k(5)*pco2*(1 - pw*pco/pco2/ph2/K2)/denom; % mol/s/kg

% FOR ADDITIONAL REACTION PATHWAYS FORMING METHANOL AND CO,
% ADD THOSE REACTION PATHWAY RATE TERMS TO r(1,1) and r(2,1) BELOW

% r(1,1) = de(1)/dW = dFmeoh/dW = rate of methanol formation
r(1,1) = R1;

% r(2,1) = de(2)/dW = dFco/dW = rate of CO formation
r(2,1) = R2;

```

## Listing of function file MeOHenergy.m

```

function dTdW = MeOHenergy(T,P,F,Bprop,r1,r2)
% by Richard K. Herz <rherz@ucsd.edu> 2011
% is called by function MeOHodes
% uses function MeOHheatrxn
% uses function MeOHheatCp
% returns dT/dW in K/kg

R = 8.314472e-05; % bar-m3/mol/K, for ideal gas law, ideal gas constant

% Bprop = [Bdiam,Btubes,Blen,Bdp,Bvoid,Bcatdens,Wcat,U,Tj];
Bdiam = Bprop(1);
Btubes = Bprop(2);
Blen = Bprop(3);
Wcat = Bprop(7);
U = Bprop(8); % J/s/m2/K, heat transfer coefficient
Tj = Bprop(9); % K, heat transfer jacket temperature

Aw = Btubes*pi*Bdiam*Blen/Wcat; % m2/kg, heat xfer area based on ID

% need total molar flow rate Ftot below
Ftot = sum(F); % mol/s

Cp = MeOHheatCp(T,P,F); % (J/K/mol), mixture heat capacity

```

```

% dH1 = delta-H of Independent Stoichiometric Eqn 1 per mol extent 1
% dH2 = delta-H of Independent Stoichiometric Eqn 2 per mol extent 2
dH = MeOHheatrxn(T,P); % J/mol
dH1 = dH(1); % J/mol
dH2 = dH(2); % J/mol

FA = U*Aw/Ftot/Cp*(Tj-T); % K/kg, heat transfer at wall term
FB = -dH1/Ftot/Cp*rl; % K/kg, ISE 1 heat of reaction term
FC = -dH2/Ftot/Cp*r2; % K/kg, ISE 2 heat of reaction term

dTdW = FA + FB + FC; % K/kg

```

## Listing of function file MeOHpress.m

```

function dPdW = MeOHpress(T,P,F,Bprop,Fin)
% by Richard K. Herz <rherz@ucsd.edu> 2011
% is called by function MeOHodes
% uses function MeOHviscosity
% returns dP/dW in bar/kg

R = 8.314472e-05; % bar-m3/mol/K

% Ergun from Fogler
% dP/dz = (-G/rho/gc/Dp)*((1-eps)/eps^3)*[150*(1-eps)*mu/Dp + 1.75*G]
% and finally
% dp/dW = dp/dz * Blen/Wcat

mu = MeOHviscosity(T,P,F); % (Pa s) = kg/m/s

% cross-sectional area of all tubes in bed
% Bprop = [Bdiam,Btubes,Blen,Bdp,Bvoid,Bcatdens,Wcat,U,Tj];
Bdiam = Bprop(1);
Btubes = Bprop(2);
Ax = Btubes*pi*Bdiam^2/4;

% gas volumetric flow rate assuming ideal gas law
Ftot = sum(F); % mol/s, total molar flow rate
flow = Ftot*R*T/P; % m3/s at this point in reactor

% molar flow rates in order: H2, CO, CO2, methanol, water, N2, CH4
MW = 1e-3*[2 28 44 32 18 28 16]; % kg/mol, molecular weights
massflow = Fin .* MW; % kg/s
massflow = sum(massflow); % kg/s
rho = massflow/flow; % kg/m3, gas density at this point in reactor
G = massflow/Ax; % kg/s/m2, superficial mass velocity
gc = 1; % kg-m/s^2/N, force unit conversion factor
% Bprop = [Bdiam,Btubes,Blen,Bdp,Bvoid,Bcatdens,Wcat];
Dp = Bprop(4);
vf = Bprop(5); % vf = void fraction, eps is std function name

dPdZ = -G/rho/gc/Dp*(1-vf)/vf^3*(150*(1-vf)*mu/Dp + 1.75*G); % Pa/m

Blen = Bprop(3);
Wcat = Bprop(7);
dPdW = dPdZ * Blen/Wcat; % Pa/kg = Pa/m * m/kg

dPdW = dPdW * 1e-5; % bar/kg = Pa/kg * 1 bar / 1e5 Pa

```

## Listing of function file MeOHmolarflow.m

```

function F = MeOHmolarflow(e,Fin)
% by Richard K. Herz <rherz@ucsd.edu> 2011
% is called by function MeOHodes
% returns current vector of molar flow rates (mol/s)

```

```
% INDEPENDENT STOICHIOMETRIC EQUATIONS (ISE) DEFINING EXTENTS e()
%   ISE 1 for e(1): CO2 + 3H2 = CH3OH + H2O
%   ISE 2 for e(2): CO2 + H2  = CO  + H2O

% molar flow rates in order: H2, CO, CO2, methanol, water, inert, total
F(1) = Fin(1) - 3*e(1) - e(2); % H2
F(2) = Fin(2)      + e(2); % CO
F(3) = Fin(3)      - e(1) - e(2); % CO2
F(4) = Fin(4)      + e(1); % methanol
F(5) = Fin(5)      + e(1) + e(2); % water
F(6) = Fin(6); % N2 (inert)
F(7) = Fin(7); % CH4 (inert)
```